



**2019/TDC/ODD/SEM/
BTCHCC-303T/243**

TDC (CBCS) Odd Semester Exam., 2019

BIOTECHNOLOGY

(3rd Semester)

Course No. : BTCHCC-303T

(Chemistry—I)

Full Marks : 50

Pass Marks : 20

Time : 3 hours

*The figures in the margin indicate full marks
for the questions*

UNIT—I

1. Answer any two of the following questions :

2×2=4

(a) State the de Broglie equation.

(b) What is Aufbau's principle?

(c) What are s-block elements?

2. (a) Describe Heisenberg's uncertainty principle and its significance.

6

Or

(b) Explain the variation of orbital energy with atomic number.



(2)

UNIT—II

3. Answer any *two* of the following questions : $2 \times 2 = 4$

- (a) Define orbital energy.
- (b) Name the elements having the smallest and largest atomic radii respectively.
- (c) Differentiate between atomic number and atomic mass.

4. (a) Differentiate between *d*- and *f*-block elements citing examples of each. $3 + 3 = 6$

Or

- (b) Write short notes on shielding effect and effective nuclear charge. $3 + 3 = 6$

UNIT—III

5. Answer any *two* of the following questions : $2 \times 2 = 4$

- (a) Define viscosity.
- (b) What do you mean by boiling point of a liquid?
- (c) Which liquid has the highest surface tension?

20J/1269

(Continued)

(3)

6. (a) What is surface tension? Explain the effect of addition of various solutes on surface tension $2 + 4 = 6$

Or

(b) Write a note on the physical properties of liquids. 6

UNIT—IV

7. Answer any *two* of the following questions : $2 \times 2 = 4$

- (a) What types of bonds are formed between molecules in water?
- (b) Define polarity.
- (c) What is detergent?

8. (a) Explain the cleansing action of detergents. 6

Or

(b) Explain how the viscosities of liquids and gases vary with a change in temperature.

UNIT—V

9. Answer any *two* of the following questions : $2 \times 2 = 4$

- (a) Define an electrolyte citing a suitable example.

20J/1269

(Turn Over)



(4)

(b) Define pH of a solution.

(c) Define pK_a of an acid.

10. (a) Discuss the common ion effect. Give one example each of a strong electrolyte and a weak electrolyte. $4+1+1=6$

Or

(b) Discuss acid-base titration curves. What is the colour of phenolphthalein in acidic and basic conditions? $4+1+1=6$

★ ★ ★